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Enabling Delta Life

Biogeochemical and hydrological controls on P retention

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## Phosphorus in surface water

#### **Phosphorus in Hunze**

Data WSHA



P-total in unfiltered samples: common practice in surface water monitoring



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## Phosphorus in surface water

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**Observation 2:** 

low dissolved concentrations

Particulate P inorganic P?

organic P?

## P speciation in surface water: samping in three areas:





#### Hunze catchment



#### Haarlemmermeer polder



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Van der Grift et al., 2015; in prep.

## Samping in three areas: particulate P and dissolved P



0.3 0.2 0.1

0.0

PLS1 PLS2 SLT1 SLT2 SLT3 SLT3 SLT4 SLT4 SLT5 XNP1 STW1 STW3

STW2

STW5 STW6

STW4

STW8

XSLT1 NPY1

STW7

PLS5

PLS3 PLS4





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## **Suspended sediment samples**

#### Hunze

#### Haarlemmermeer

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### Sequential extraction: P fractionation in suspended sediment



#### Hunze catchment



#### Haarlemmermeer



## Why dominance of Fe bound P?

Lowland area: reactive subsurface with high organic Iron in shallow groundwater (8-10 m depth) matter content

- anaerobic groundwater
- high iron(II) concentration
- contributes to surface water

So, high load of iron(II) to surface water system

- 1: Oxidation iron(II) with oxygen to Fe hydroxides  $Fe^{2+} + 0.25 O_2 + 2.5 H_2O \rightarrow Fe(OH)_3 + 2H^+$
- 2: Phosphate: strong affinity with Fe hydroxides:

Surface adsorption on Fe-hydroxides surfaces  $\equiv XOH_2^+ + HPO_4^{2-} \rightarrow \equiv XOH_2 HPO_3^- + H_2O$ 

Precipitation Fe phosphate phases  $Fe(OH)_3 + PO_4^{3-} \rightarrow FePO_4 + 3OH$ 







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# Fe and P dynamics in at the groundwater – surface water interface

## Hupsel brook field experiment



Dissolved total P concentrations in Hupsel catchment





Van der Grift et al., 2014; HESS.

## Fe<sup>2+</sup> oxidation experiments in presence of PO



Reactor to oxidize Fe<sup>2+</sup> under controlled conditions:

- Fixed pH
- Fixed O<sub>2</sub> concentrations

P/Fe initial ratio = 0.18 pH=6.1 & O<sub>2</sub>=10.5mg/l 200 1.0  $Fe^{2+}$ from NaOH addition;  $H^+/Fe^{2+} = 1.55$  $Fe^{2+}$ from NaOH addition;  $H^+/Fe^{2+} = 1.89$ 0.9 Fe<sup>2+</sup> extraction 160 0.8 PO<sub>4</sub> from NaOH & solid P/Fe = 0.41 Fe(II) & PO<sub>4</sub> (µmol/l) 8 05 PO<sub>4</sub> extraction 0.7 P/Fe ratio precipitates D/Fe ratio P/Fe ratio 0.3 0.2 40 ← PO4 depletion 0.1 -0.0 0 1000 500 1500 2000 2500 3000 time (min)

PO<sub>4</sub> depleted:

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Fast binding of  $PO_4$  during oxidation of Fe<sup>2+</sup>:

- No time for dissolved P transport

PO<sub>4</sub> present:

## Conceptual model of abiotic in-stream processes influencing phosphate transport



Inspiration from continuous water quality / sediment measurements

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# Effect of flow velocities on re-suspension bed sediments and particulate P in polder catchment

10 minute interval measurements

- Suspended sediment (Turbidity)
- Total-P and total reactive P
- NO<sub>3</sub>
- conductivity





#### Blocq van Kuffeler pumping station



## High-Frequent monitoring at pumping-station in Flevoland

#### October 2014 – April 2015





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## Response on pumping: increase of TP

Total-P



## Total P remoblisation: natural catchment vs polder



## **Conclusions & questions left**

- Fe-bound P is dominant P species in surface water in agricultural lowland areas
- Seepage of anaerobic groundwater controls P speciation in surface water
- Fast immobilization of P during oxidation process of Fe(II)
  - Precipitation of Fe hydroxyphosphate phase
  - How stable are these mineral phases?
  - Bioavailability?
- Flow induced resuspension of streambed sediment is main transport mechanism for P
  - Are natural catchments much more vulnerable for incidental losses than polder??